***ASSIGNMENT***

1. ***Calculate the quantity of electricity required to deposit/erode one mole of copper at the cathode/anode(Cu=63.5)***
2. ***Determine the Faraday constant from the results in (i) above.***
3. ***Name the products of electrolysis of 1M copper(II) sulphate(VI) solution using copper electrodes.***
4. ***During purification of copper by electrolysis 1.48 g of copper was deposited when a current was passed through aqueous copper (II)sulphate(VI) for 2 ½ hours. Calculate the amount of current that was passed. (Cu= 63.5 , 1F = 96500C)***
5. ***An element “P” has a relative atomic mass 88.When a current of 0.5 amperes was passed through fused chloride of “P” for 32 minutes and 10seconds ,0.44 g of “P” was deposited at the cathode. Determine the charge on an ion of “P”(Faraday constant = 96500C)***